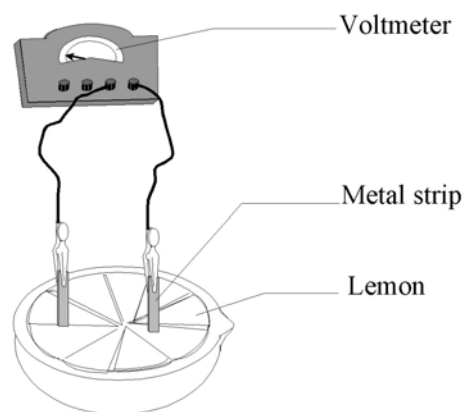


Answer questions 5-16 by referring to Table J. For each of the electrode pairs, which would be the anode in an electrochemical cell?

- | | | |
|----------------|-----------------|------------------------------|
| 5. Cu/Zn | 9. Au/Pb | 13. Co/Ni |
| 6. Pb/Sn | 10. Mn/Zn | 14. H ₂ /Ag |
| 7. K/Al | 11. Fe/Zn | 15. Cu/Mg |
| 8. Ba/Li | 12. Co/Ca | 16. Zn/Al |

Answer questions 17-19 by referring to the setup shown to the right using a lemon and metal strips. It actually produces measurable electricity.



17. Explain how the lemon battery works? _____

18. What parts of a typical voltaic cell are missing in the lemon battery? What effect does this have on how well it functions? Explain. _____

19. If the metal strip on the right is iron and the metal strip on the left is aluminum, in what direction will electricity flow? _____



20. What happens at the anode of an electrochemical cell? _____

21. There are two voltaic cells pictured on the previous page. The one on the left is called a wet cell, while the one at the left is called a dry cell. The one at the right is also called an alkaline cell. What is the difference between these cells that accounts for the difference in the way they are named? _____

