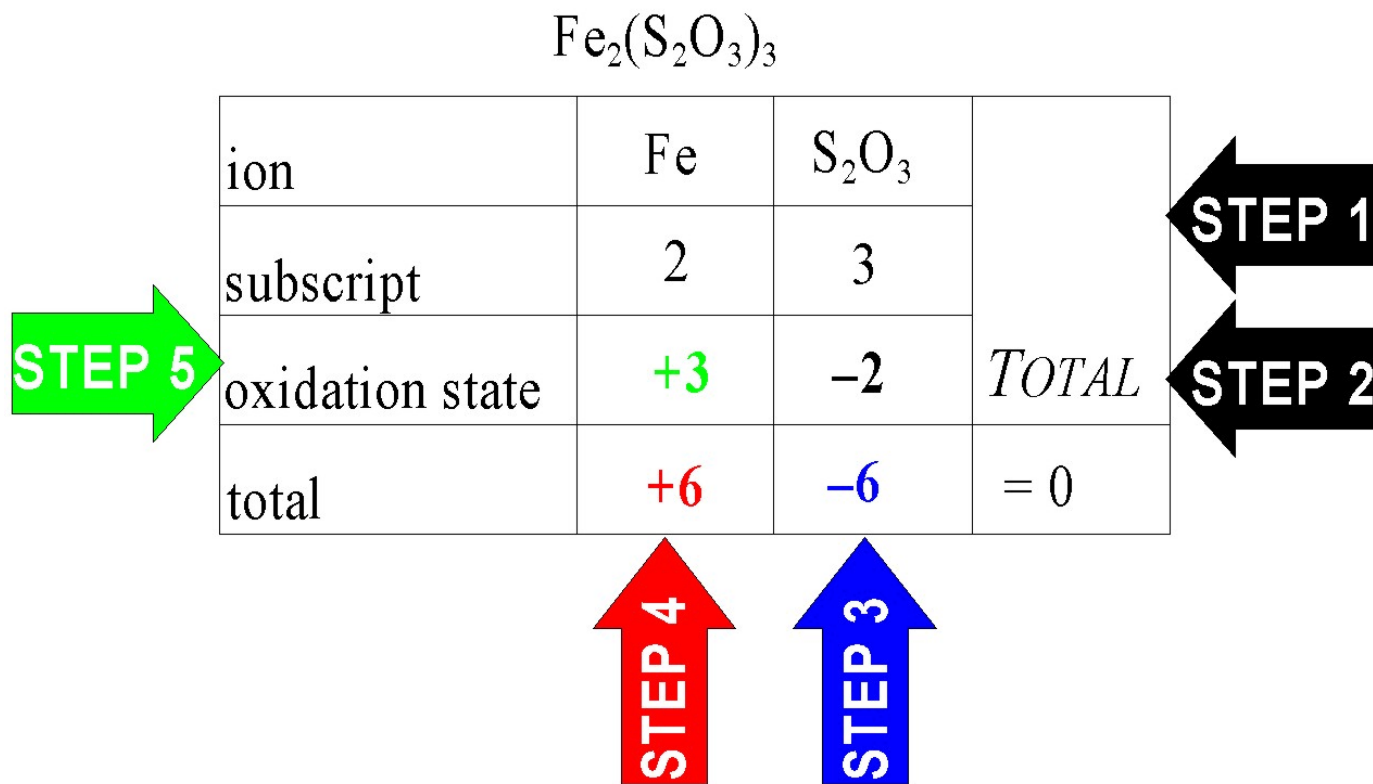


Determining the Charge on a Metal Ion

$\text{Fe}_2(\text{S}_2\text{O}_3)_3$

ion	Fe	S_2O_3	
subscript	2	3	
oxidation state	+3	-2	TOTAL
total	+6	-6	= 0



Prepare a table as above.

Step 1: List the subscripts for the metal and the nonmetal ions.

Step 2: Look up the oxidation state of the nonmetal ion on the *Periodic Table*.

Step 3: Multiply the oxidation state of the nonmetal by its subscript to get the total charge.

Step 4: Determine the total charge of the metal ions by calculating the number which, when added to the total charge of the nonmetal ion, gives the compound a total charge of zero.

Step 5: Divide the total charge of the metal ions by the subscript of the metal to get the oxidation state.