Chemistry:	Form	OH9.2.1A	
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ACIDS, BASES, AND SALTS

Date _____ Period ____

Comparing Acids and Bases

Characteristic	Acids	Bases
Conductivity	 Aqueous solutions are electrolytes (Conduct electricity) ★ The ability of acids to conduct electricity is proportional to their degree of ionization ★ strong acids are strong electrolytes ★ weak acids are weak electrolytes 	Aqueous solutions are electrolytes ★ strong bases are strong electrolytes ★ weak bases are weak electrolytes
рН	Increase the hydronium ion concentration of water ★ Have a pH below 7	Increase the hydroxide ion concentration of water ★ Have a pH above 7
Taste	Taste sour	Taste bitter
Indicators	Cause color changes in indicators (indicator - something that reacts with an acid or base to show a definite color change) ★ litmus → red ★ phenolphthalein → clear ★ bromthymol blue → yellow ★ methyl orange → red	Cause color changes in indicators
Typical Reaction	Corrosive – React with active metals (below hydrogen on Table J <i>Activity Series</i>) to release hydrogen (except for oxidizing acids like HNO ₃)	Caustic – Feel slippery because bases dissolve skin
Neutralization	React with bases to form a salt and water	React with acids to form a salt and water