Drawing Bohr Diagrams

Visualizing Atoms

Getting the Data

- Information about the atom comes from the *Periodic Table*.
- Example: Sodium

Mass:	22.98977 amu
Atomic Number:	11
Electron Configuration:	2–8–1

Using the Data

- Unless the mass number (A) of a specific isotope is given, round off the atomic mass to get the mass number of the most common isotope: 22.98977 amu ≈ 23 amu
- You should now have the following data:

A	Z	N	Electron Configuration
23	11	12	2–8–1

- Subtract the atomic number from the mass number to get the number of neutrons.

Drawing from the Data

A	Z	N	Electron Configuration
23	11	12	2–8–1

- Write the number of protons followed by “P.”
- Write the number of neutrons followed by “N.”
- Draw an arc to represent each energy level with the correct number of electrons followed by the symbol “e^-.”