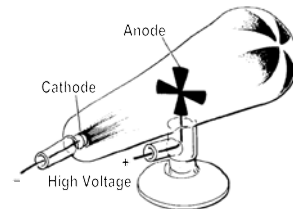


Test Review No 4

Dalton's Atomic Theory. Dalton proposed atomic theory in 1803 to explain his observations about the relative masses of elements in a compound. Dalton's Postulates say: [1] Matter is made of small particles called atoms; [2] Atoms are indestructible. They cannot be created or destroyed during chemical or physical changes; [3] Atoms of an element are identical. They have the same mass; [4] Atoms of different elements have different masses; [5] Compounds are formed by combining atoms of different elements. The Dalton model of the atom is a solid, indivisible sphere.

Thomson's Model. Thomson showed that the beam of light in a Crookes tube was actually composed of negatively charged particles he called electrons. In order to account for the negatively charged particles in neutral matter, he assumed the rest of the atom was positively charged. Thomson's model of the atom was a positively charged cloud with the negative electrons scattered evenly throughout.

Rutherford's Model. Ernest Rutherford performed the alpha scattering experiment in 1911. He probed the inside of the atom by aiming small, positively charged particles called alpha particles at gold foil. Only 1 in 8,000 alpha particles bounced straight back or were deflected greatly. The rest went straight through the gold foil. Based on these observations, Rutherford suggested that the atom is mostly empty space with a small, positively charged center and negatively charged electrons revolving around the outside like planets around the sun.



The Bohr Model. Bohr developed a model of the atom with circular pathways for the electron. These pathways were at fixed distances from the nucleus. Electrons could be found only in these circular pathways. If an electron absorbed enough energy, it could jump up to another level, but it could never be found between levels. Inevitably, the electron lost energy and fell back down to a previous level, giving off the extra energy as a specific frequency of light. Bohr had a complex equation into which he could substitute simple numbers, integers such as 1, 2, or 3, and the equation predicted the frequencies of the light. These mystery numbers represented the energy levels of the electrons. Bohr's evidence for his model was the bright line spectra formed when electrons fell from the excited state back to the ground state.

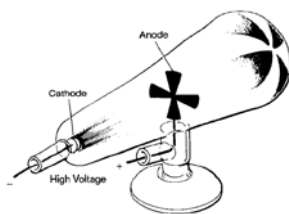
Wave Mechanical Model. The Bohr model successfully explained the bright line spectra for hydrogen, but could not explain the spectra of atoms with more electrons. The wave mechanical model solved the problem. The wave mechanical model describes the location of electrons as their most probable location rather than as orbits with fixed radii. The regions where electrons are most probably found are called orbitals. An orbital can hold, at most, two electrons.

Answer the questions below by circling the number of the correct response

- In 1661, Robert Boyle proposed that matter is composed of simple substances called elements that cannot be broken down by chemical means. His theory was not accepted for over 100 years because (1) he lacked qualitative evidence, (2) he lacked quantitative evidence, (3) he lacked both qualitative and quantitative evidence, (4) he was a philosopher rather than a scientist.
- Boyle proposed that matter is composed of simple substances called elements that cannot be further decomposed or broken down. Aristotle's idea that air was an element did not fit Boyle's definition because air (1) cannot be seen, (2) cannot be broken down, (3) can be decomposed, (4) can react with other substances.
- Antoine Lavoisier measured the mass of some mercury. Then he heated it in air. The mercury turned into a red substance as it was heated. The mass of the red substance was greater than the mass of the mercury. The best explanation for this observation is that (1) matter was created, (2) a new element formed, (3) the mercury combined with something in the air, (4) hot mercury is heavier than cold mercury.
- Modern atomic theory is based on the work of (1) Kepler, (2) Aristotle, (3) Dalton, (4) Leeuwenhoek.
- Lavoisier heated mercury until a red substance formed. Then he heated the red substance and it broke down into mercury and oxygen. From his observations it is reasonable to conclude (1) mercury and oxygen are elements and the red substance is a compound composed of mercury and oxygen, (2) mercury, oxygen, and the red substance are elements, (3) the red substance is an element, and mercury and oxygen are compounds composed from it, (4) mercury, oxygen, and the red substance are compounds.
- When Lavoisier heated mercury II oxide, it decomposed into mercury and oxygen. The mass of the mercury II oxide was equal to the mass of the mercury plus the mass of the oxygen. This shows that during the chemical reaction (1) matter was not created but it was destroyed, (2) matter was created but it was not destroyed, (3) matter was both created and destroyed, (4) matter was neither created nor destroyed.
- An 18 kg sample of water is decomposed by electrolysis, releasing 16 kg of oxygen. How much hydrogen was released? (1) 34 kg (2) 2 kg (3) 16 kg (4) 1.125 kg
- If 46 g of X combines with 16 g of Y to form Z, how much Z is formed? (1) 30 g (2) 2.9 g (3) 724 g (4) 62 g

REVIEW

9. When William Crookes passed current through a cathode (negative electrode) in a glass vacuum tube, the tube glowed casting a shadow of the anode at the opposite end as shown in the diagram at the right. This showed Crookes that the source of the light was the (1) anode, (2) cathode, (3) protons, (4) neutrons.



10. Which particles have the smallest mass? (1) electrons (2) protons (3) neutrons (4) atoms
11. Which of the following would *NOT* help to explain Dalton's observations about the relative masses of elements in a compound. (1) Matter is made of small particles called atoms. (2) Atoms are divisible. (3) Atoms of an element have the same mass. (4) Compounds are formed by combining atoms of different elements.
12. Dalton's model of the atom was (1) a solid sphere, (2) a positive cloud with scattered electrons, (3) the "solar system" model, (4) based on electrons moving in fixed circular pathways around the nucleus
13. Who made the discovery that cathode rays were actually negatively charged particles called electrons? (1) Thomson, (2) Bohr, (3) Rutherford, (4) Dalton
14. Whose model of the atom could be represented by a solid, indivisible sphere with a characteristic mass? (1) Dalton (2) Thomson (3) Rutherford (4) Bohr
15. Whose model of the atom could be represented by a positively charged cloud with electrons distributed through it? (1) Dalton (2) Thomson (3) Rutherford (4) Bohr
16. The nucleus of the atom was discovered by (1) Thomson, (2) Bohr, (3) Rutherford, (4) Dalton
17. Subatomic particles can usually pass undeflected through an atom because the volume of an atom is composed of (1) an uncharged nucleus (2) largely empty space (3) neutrons (4) protons
18. Which of the following particles is negatively charged? (1) electron (2) proton (3) neutron (4) cation

19. Whose model of the atom could be represented by the diagram to the right? (1) Dalton (2) Thomson (3) Rutherford (4) Bohr



20. Evidence that electrons exist in distinct energy levels outside the nucleus is provided by (1) cathode rays, (2) spectral lines, (3) atomic masses, (4) radioactivity.

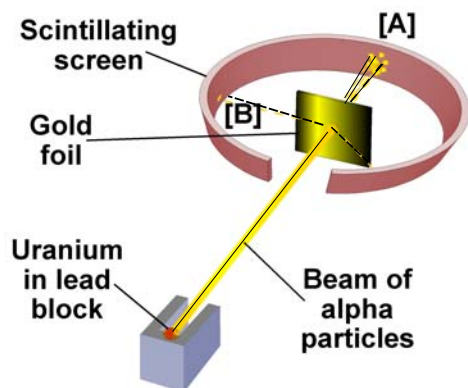
21. When excited electrons return to the ground state, they lose excess energy in the form of (1) light, (2) gamma rays, (3) nuclear radiation, (4) sound.
22. When the electrons of an excited atom return to a lower energy state, the energy emitted can result in the production of (1) alpha particles (2) isotopes (3) protons (4) spectra
23. Which electron transition represents a gain of energy? (1) from 2nd to 3rd shell (2) from 2nd to 1st shell (3) from 3rd to 2nd shell (4) from 3rd to 1st shell
24. Which statement best describes electrons? (1) They are positive subatomic particles and are found in the nucleus. (2) They are positive subatomic particles and are found surrounding the nucleus. (3) They are negative subatomic particles and are found in the nucleus. (4) They are negative subatomic particles and are found surrounding the nucleus.
25. During a flame test, ions of a specific metal are heated in the flame of a gas burner. A characteristic color of light is emitted by these ions in the flame when the electrons (1) gain energy as they return to lower energy levels (2) gain energy as they move to higher energy levels (3) emit energy as they return to lower energy levels (4) emit energy as they move to higher energy levels
26. Which of the following particles is negatively charged? (1) electron (2) proton (3) neutron (4) cation
27. The modern model of the atom shows that electrons are (1) orbiting the nucleus in fixed paths, (2) found in regions called orbitals, (3) combined with neutrons in the nucleus, (4) located in a solid sphere covering the nucleus
28. The characteristic bright-line spectrum of an element is produced when electrons (1) fall back to lower energy levels, (2) are gained by a neutral atom, (3) are emitted by the nucleus as beta particles, (4) move to higher energy levels.
29. The Bohr model is supported by (1) the spectrum of hydrogen, but not the spectrum of other elements, (2) the spectrum of other elements, but not the spectrum of hydrogen, (3) both the spectrum of hydrogen, and the spectrum of other elements, (4) neither the spectrum of hydrogen, nor the spectrum of other elements.
30. Both the Bohr model and the quantum mechanical model agree that (1) the electrons move in fixed circular pathways, (2) the electrons can be anywhere outside the nucleus, (3) the energy of the electron is quantized, (4) the electrons closest to the nucleus have the highest energy.
31. Which of the following atomic models does NOT have moving electrons? (1) Thomson (2) Rutherford (3) Bohr (4) Quantum mechanical

REVIEW

Page 3

Answer questions 32-35 by referring to the information and diagram below.

Rutherford performed the following experiment: He aimed a beam of high speed, positively charged particles called alpha particles at a piece of solid gold foil. He set up a special screen all around the foil to help him see where the particles went.



32. What would have happened to the alpha particles when they hit the gold foil if Dalton's model of the atom were correct? (1) All the alpha particles would have passed through. (2) All the alpha particles would have bounced back. (3) Most of the alpha particles would have passed through, but some would have bounced back. (4) Most of the alpha particles would have bounced back, but some would have passed through.
33. What would have happened to the alpha particles when they hit the gold foil if Thomson's model of the atom were correct? (1) All the alpha particles would have passed through. (2) All the alpha particles would have bounced back. (3) Most of the alpha particles would have passed through, but some would have bounced back. (4) Most of the alpha particles would have bounced back, but some would have passed through.
34. When Rutherford performed his experiment, only 1 in 8,000 alpha particles bounced straight back or were deflected greatly [B]. The rest went straight through the gold foil. [A] Based on this evidence, what is in an atom's center? (1) electrons (2) a small, massive, positively charged nucleus (3) a large, massive, negatively charged nucleus (4) empty space
35. Based on his evidence, what model of the atom did Rutherford propose? (1) The atom is a solid indivisible ball. (2) The atom is a positively charged cloud with electrons distributed evenly throughout. (3) The atom has a positively charged nucleus with electrons revolving around it like planets around the sun. (4) The energy of the electrons is quantized and they are found in orbitals.

35.	3	28.	1	21.	1	14.	1	7.	2
34.	2	27.	2	20.	2	13.	1	6.	4
33.	1	26.	1	19.	4	12.	1	5.	1
32.	2	25.	3	18.	1	11.	2	4.	3
31.	1	24.	4	17.	2	10.	1	3.	3
30.	3	23.	1	16.	3	9.	2	2.	3
29.	1	22.	4	15.	2	8.	4	1.	2

Answers